

RARE EARTH COMPOSITIONS AND STRUCTURES FOR REMOVING PHOSPHATES FROM WATER

5 **[0001]** The present invention is a continuation-in-part application of USSN 60/430,284 filed December 2, 2002, the entire contents of which are incorporated herein by reference.

10 **[0002]** The present invention relates to a chemical composition and a physical structure of a chemical compound, used to efficiently remove phosphates from water. Particularly, the invention relates to the use of rare-earth compounds to control algal growth in swimming pools and other water systems. More particularly, the invention relates to lanthanum compounds. The description of the invention is based on the use of lanthanum. It is to be understood that other rare-earth elements can be substituted for lanthanum.

BACKGROUND OF THE INVENTION

15 **[0003]** Traditional algal control in swimming pools and other water systems is achieved by biocides. This generally requires substantial amounts of toxic chemicals.

20 **[0004]** New methods that have recently been developed are based on the removal of phosphate, an indispensable nutrient for algal growth, from the water. Several methods and compositions based on lanthanum compounds have recently been proposed for the removal of phosphate from water to control algal growth. US Patent 6,146,539 discloses a treatment method for swimming pool water based on the addition of finely divided, insoluble, lanthanum carbonate or of soluble lanthanum chloride. The lanthanum carbonate reaction is typically slow, and several days are required to see an effect in practice. Lanthanum chloride produces a milky precipitate that can only be removed by the addition of copious amounts of flocculent. Making styrene-based ion-exchange beads incorporating lanthanum carbonate was also effective but slow: in one example, it took 4 days to reduce the phosphate concentration from 400 to 30 ppb.

30 **[0005]** US Patent 6,312,604 uses a polymer e.g. polyacrylamide or polyvinyl alcohol with a binder to attach a lanthanide halide salt, preferably La chloride. This method prevents the formation of very fine precipitate, but the reaction rates are also very slow.

[0006] A method that has been proposed to accelerate the rate of formation of lanthanum phosphate is to use a compound with intermediate solubility, such as lanthanum sulfate, either alone or in combination with La carbonate. Such method is disclosed in US Patent 6,338,800. One drawback of the method is that
5 excess lanthanum sulfate will leave lanthanum in solution.

[0007] Lanthanum oxycarbonates have recently been disclosed to remove phosphate from the gastro-intestinal tract and the bloodstream in patients with hyperphosphatemia. We have now found that the properties of lanthanum oxycarbonates can also be applied to efficiently remove phosphates from water to
10 very low levels.

SUMMARY OF THE INVENTION

[0008] In accordance with the present invention, rare-earth compounds, and in particular, rare earth oxycarbonates are provided. The oxycarbonates may be hydrated or anhydrous. These compounds may be produced according to the
15 present invention as particles having a porous structure. The rare-earth compound particles of the present invention may conveniently be produced in a controllable range of surface areas with resultant variable and controllable adsorption or chemical reaction rates of the phosphate ion.

[0009] It has now been found that the properties of lanthanum
20 oxycarbonate can provide unexpected advantages over lanthanum carbonate, lanthanum halides (particularly chloride) and lanthanum sulfate for the removal of phosphate from water for the prevention of algal growth. The lanthanum compounds of this invention are lanthanum oxycarbonates, particularly $\text{La}_2\text{O}(\text{CO}_3)_3 \cdot 4\text{H}_2\text{O}$ and $\text{La}_2\text{O}_2\text{CO}_3$. These compounds can be made by any method.

[0010] A method of making lanthanum oxycarbonate hydrate particles
25 includes making a solution of lanthanum chloride, subjecting the solution to a slow, steady feed of a sodium carbonate solution at a temperature between about 30°C and about 90°C while mixing, then filtering and washing the precipitate, then drying the filter cake at a temperature between about 100°C and about 120°C to produce
30 the desired lanthanum oxycarbonate hydrate species. Optionally, the filter cake may be dried then slurried and milled in a horizontal or vertical pressure media mill to a

desired surface area, spray dried or dried by other means to produce a powder that may be washed, filtered and dried.

[0011] Another method of making the anhydrous lanthanum oxycarbonate particles includes making a solution of lanthanum chloride, subjecting the solution to a slow, steady feed of a sodium carbonate solution at a temperature of about 30°C to 90°C while mixing, then filtering and washing the precipitate, then drying the filter cake at a temperature between about 100°C and 120°C to produce the desired lanthanum oxycarbonate hydrate species. Then the dried filter cake is subjected to a thermal treatment at a temperature between 400°C to 700°C. Optionally, the product of the thermal treatment may be slurried and milled in a horizontal or vertical pressure media mill to a desired surface area, spray dried or dried by another means to produce a powder that may be washed, filtered and dried.

[0012] Still another method of making anhydrous lanthanum oxycarbonate particles includes making a solution of lanthanum acetate, subjecting the solution to a total evaporation process using a spray dryer or other suitable equipment to make an intermediate product, and calcining the intermediate product obtained at a temperature between about 500° and about 1200°C. The intermediate product of the calcination step may be washed, filtered and dried to make a suitable final product. Optionally the intermediate product may be milled in a horizontal or vertical pressure media mill to a desired surface area, spray dried or dried by other means to produce a powder that may be washed, filtered and dried.

[0013] The porous particles or porous structures of the present invention are made of nano-sized to micron-sized crystals. The lanthanum oxycarbonate hydrate is preferably lanthanum oxycarbonate tri or tetra hydrate $(\text{La}_2\text{O}(\text{CO}_3)_2 \cdot x\text{H}_2\text{O})$ where $2 \leq x \leq 4$, including where x is 3 or 4. The preferred anhydrous lanthanum oxycarbonate is $\text{La}_2\text{O}_2\text{CO}_3$, also written as $(\text{LaO})_2\text{CO}_3$ or La_2CO_5 , of which several crystalline forms exist.

BRIEF DESCRIPTION OF THE DRAWINGS

[0014] FIG. 1 is a general flowsheet of a process according to the present invention that produces lanthanum oxycarbonate tri or tetra hydrate $(\text{La}_2\text{O}(\text{CO}_3)_2 \cdot x\text{H}_2\text{O})$, with where $2 \leq x \leq 4$, including where x is 3 or 4.

[0015] FIG. 2 is a scanning electron micrograph of a lanthanum oxycarbonate $\text{La}_2\text{O}(\text{CO}_3)_2 \cdot x\text{H}_2\text{O}$ (where $2 \leq x \leq 4$, including where x is 3 or 4) porous structure made according to the process of the present invention and magnified 120,000 fold.

5 [0016] FIG. 3 is an XRD scan of lanthanum oxycarbonate hydrate ($\text{La}_2\text{O}(\text{CO}_3)_2 \cdot x\text{H}_2\text{O}$) wherein where $2 \leq x \leq 4$, including where x is 3 or 4 and generally close to 4 and wherein the lanthanum oxycarbonate hydrate is made according to the process of the present invention and compared with a standard library card of $\text{La}_2\text{O}(\text{CO}_3)_2 \cdot x\text{H}_2\text{O}$.

10 [0017] FIG. 4 is a general flow sheet of a process according to the present invention that produces anhydrous lanthanum oxycarbonate ($(\text{LaO})_2\text{CO}_3$ or La_2CO_5 , of which several crystalline forms exist).

[0018] FIG. 5 is a scanning electron micrograph of lanthanum oxycarbonate ($(\text{LaO})_2\text{CO}_3$ or La_2CO_5 , of which several crystalline forms exist) porous structure made according to the process of the present invention and magnified 60,000 fold.

[0019] FIG. 6 is an XRD scan of anhydrous lanthanum oxycarbonate ($(\text{LaO})_2\text{CO}_3$ or La_2CO_5 , of which several crystalline forms exist) made according to the process of the present invention and compared with a standard library card of $\text{La}_2\text{O}_2\text{CO}_3$. The bottom of the figure shows another phase of lanthanum oxycarbonate La_2CO_5 .

[0020] FIG. 7 is a general flow sheet of a process according to the present invention that produces anhydrous lanthanum oxy-carbonate ($(\text{LaO})_2\text{CO}_3$ or La_2CO_5 of which several crystalline forms exist).

25 [0021] FIG. 8 is a scanning electron micrograph of lanthanum oxycarbonate ($(\text{LaO})_2\text{CO}_3$ or La_2CO_5 or which several crystalline forms exist) porous structure, magnified 80,000 fold.

[0022] FIG. 9 is an XRD scan of lanthanum oxycarbonate ($(\text{LaO})_2\text{CO}_3$ or La_2CO_5 of which several crystalline forms exist) as produced and compared with a standard library card of lanthanum oxy-carbonate ($\text{La}_2\text{O}_2\text{CO}_3$). The bottom of the figure shows another phase of La_2CO_5 (Lanthanum oxycarbonate).

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[0023] FIG. 10 is a graph comparing the reaction rate of commercial grades of lanthanum carbonate ($\text{La}_2(\text{CO}_3)_3 \cdot 4\text{H}_2\text{O}$ and $\text{La}_2(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$), with the reaction rates of the lanthanum oxycarbonate tetra hydrate and the anhydrous oxycarbonates of this invention.

5 DESCRIPTION OF THE INVENTION

[0024] Referring now to the drawings, the process of the present invention will be described. While the description will generally refer to lanthanum compounds, the use of lanthanum is merely for ease of description and is not intended to limit the invention and claims solely to lanthanum compounds. In fact, it is contemplated that
10 the process and the compounds described in the recent specification are equally applicable to lanthanides and rare earth metals other than lanthanum, such as Ce and Y.

[0025] Turning now to FIG. 1, a process for making lanthanum oxycarbonate and in particular, lanthanum oxycarbonate tetrahydrate, is shown.
15 First, an aqueous solution of lanthanum chloride is made by any method. One method to make the solution is to dissolve commercial lanthanum chloride crystals in water or in an HCl solution. Another method to make the lanthanum chloride solution is to dissolve lanthanum oxide in a hydrochloric acid solution.

[0026] The LaCl_3 solution is placed in a well-stirred tank reactor. The LaCl_3
20 solution is then heated to a temperature between 30°C and 90°C . A previously prepared analytical grade sodium carbonate is steadily added with vigorous mixing. The mass of sodium carbonate required is calculated at 6 moles of sodium carbonate per 2 moles of LaCl_3 . When the required mass of sodium carbonate solution is added the resultant slurry or suspension is allowed to cure for about 2
25 hours at 30 to 90°C . The suspension is then filtered and washed with demineralized water to produce a clear filtrate. The filter cake is placed in a convection oven at 100 to 120°C for 1 to 5 h or until a stable weight is observed. The initial pH of the LaCl_3 solution is 2, while the final pH of the suspension after cure is 5.5. A white powder is produced. The resultant powder is a lanthanum oxycarbonate hydrate
30 $(\text{La}_2\text{O}(\text{CO}_3)_2 \cdot x\text{H}_2\text{O})$ where $2 \leq x \leq 4$, including where x is 3 or 4.

EXAMPLE I

[0027] An aqueous solution having a volume of 335 ml and containing lanthanum chloride (LaCl_3) at a concentration of 29.2 weight% as La_2O_3 was added to a 4-liter beaker and heated to 80°C with stirring. The initial pH of the LaCl_3 solution was 2.2. A volume of 265 ml of an aqueous solution containing 63.6 g of sodium carbonate (Na_2CO_3) was metered into the heated beaker using a small pump at a steady flow rate for 2 h. Using a Buchner filter apparatus fitted with filter paper, the filtrate was separated from the white powder product. The filter cake was mixed 4 times with 2 liters of distilled water and filtered to wash away the NaCl formed during the reaction. The washed filter cake was placed into a convection oven set at 105°C for 2 h or until a stable weight was observed. FIG. 2 shows a scanning electron micrograph of the product, enlarged 120,000 times. The X-Ray diffraction pattern of the product (FIG. 3) shows that it consists of hydrated lanthanum oxycarbonate $\text{La}_2\text{O}(\text{CO}_3)_2 \cdot x\text{H}_2\text{O}$, with where $2 \leq x \leq 4$, including where x is 3 or 4. The sample has a surface area measured by the BET method, of $38.8 \text{ m}^2/\text{g}$.

[0028] Turning now to FIG. 4, a process for making anhydrous lanthanum oxycarbonate is shown. First, an aqueous solution of lanthanum chloride is made by any method. One method to make the solution is to dissolve commercial lanthanum chloride crystals in water or in an HCl solution. Another method to make the lanthanum chloride solution is to dissolve lanthanum oxide in a hydrochloric acid solution.

[0029] The LaCl_3 solution is placed in a well-stirred tank reactor. The LaCl_3 solution is then heated to a temperature between 30 and 90°C . A previously prepared analytical grade sodium carbonate is steadily added with vigorous mixing. The mass of sodium carbonate required is calculated at 6 moles of sodium carbonate per 2 moles of LaCl_3 . When the required mass of sodium carbonate solution is added the resultant slurry or suspension is allowed to cure at 30 to 90°C . The suspension is then washed and filtered removing NaCl (a byproduct of the reaction) to produce a clear filtrate. The filter cake is placed in a convection oven at 100 to 120°C for 1 to 5 hours or until a stable weight is observed. The initial pH of the LaCl_3 solution is 2.2, while the final pH of the suspension after cure is 5.5. A white lanthanum oxycarbonate tetra hydrate powder is produced. Next the

lanthanum oxycarbonate tetra hydrate is placed in an alumina tray, which is placed in a high temperature muffle furnace. The white powder is heated to 400 to 700°C and held at that temperature for 2 to 5 hours. Anhydrous La_2CO_3 is formed. The compound is also designated $\text{La}_2\text{O}_2\text{CO}_3$ or $(\text{LaO})_2\text{CO}_3$.

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EXAMPLE II

[0030] An aqueous solution having a volume of 335 ml and containing lanthanum chloride (LaCl_3) at a concentration of 29.2 weight% as La_2O_3 was added to a 4-liter beaker and heated to 80°C with stirring. The initial pH of the LaCl_3 solution was 2.2. A volume of 265 ml of an aqueous solution containing 63.6 g of sodium carbonate (Na_2CO_3) was metered into the heated beaker using a small pump at a steady flow rate for 2 h. Using a Buchner filter apparatus fitted with filter paper, the filtrate was separated from the white powder product. The filter cake was mixed 4 times with 2 liters of distilled water and filtered to wash away the NaCl formed during the reaction. The washed filter cake was placed into a convection oven set at 105°C for 2 h or until a stable weight was observed. Finally, the lanthanum oxycarbonate was placed in an alumina tray in a muffle furnace. The furnace temperature was ramped to 500 °C and held at that temperature for 3h. The resultant product was determined to be anhydrous lanthanum oxycarbonate $\text{La}_2\text{O}_2\text{CO}_3$, with a surface area of 27 m²/g. FIG. 5 shows a scanning electron micrograph of the product, enlarged 60,000 times. The X-Ray diffraction pattern of the product (FIG. 6) shows that it consists of anhydrous lanthanum oxycarbonate $\text{La}_2\text{O}_2\text{CO}_3$.

[0031] Turning now to FIG. 7, another process for making anhydrous lanthanum carbonate is shown. First, a solution of lanthanum acetate is made by any method. One method to make the solution is to dissolve commercial lanthanum acetate crystals in water or in an HCl solution. Another method to make the lanthanum acetate solution is to dissolve lanthanum oxide in an acetic acid solution.

[0032] The product solution is further evaporated to form an intermediate product. The evaporation 20 is conducted under conditions to achieve substantially total evaporation. In particular, the evaporation is conducted at a temperature higher than the boiling point of the feed solution but lower than the temperature where significant crystal growth occurs. The resulting intermediate may desirably be an

amorphous solid formed as a thin film and may have a spherical shape or a shape in part of a sphere.

[0033] The term "substantially total evaporation" or "substantially complete evaporation" refers to evaporation such that the solid intermediate contains less than 15% free water, preferably less than 10% free water, and more preferably less than 1% free water. The term "free water" is understood and means water that is not chemically bound and can be removed by heating at a temperature below 150° C. After substantially total evaporation or substantially complete evaporation, the intermediate product will have no visible moisture present.

[0034] The evaporation process may be conducted in a spray dryer. In this case, the product will consist of a structure of spheres or parts of spheres. The spray dryer generally operates at a discharge temperature between about 120°C and 500°C.

[0035] The intermediate product may then be calcined 30 by raising the temperature to a temperature between about 400°C to about 800°C for a period of time from about 2 to about 24 h and then cooled to room temperature. The cooled product may be washed 40 by immersing it in water or dilute acid, to remove traces of any water-soluble phase that may still be present after the calcination step.

[0036] The temperature and the length of time of the calcination process may be varied to adjust the particle size and the reactivity of the product.

[0037] The particles obtained after calcination and washing have been used to efficiently remove phosphate from water. The particles may also be used in a device to directly remove phosphate from water.

[0038] The particles generally have a size between 1 and 1000 µm. The particles consist of individual crystals, bound together in a structure with good physical strength. They form a porous structure. The individual crystals generally have a size between 20 nm and 10 µm. If the evaporation process is conducted in a spray-dryer, the particles consist of spheres or parts of spheres.

EXAMPLE III

[0039] A solution containing 100 g/l of La as lanthanum acetate is injected in a spray dryer with an outlet temperature of 250°C. The intermediate product

corresponding to the spray-drying step is recovered in a bag filter. This intermediate product is calcined at 600 °C for 4 h. FIG. 8 shows a scanning electron micrograph of the product, enlarged 60,000 times. The X-Ray diffraction pattern of the product (FIG. 9) shows that it consists of anhydrous lanthanum oxycarbonate La_2CO_5 . The surface area of the sample, measured by the BET method, was 25 m²/g.

EXAMPLE IV:

[0040] To determine the reactivity of the lanthanum compounds with respect to phosphate, the following tests were conducted. A stock solution containing 13.75 g/l of anhydrous Na_2HPO_4 and 8.5 g/l HCl was prepared. The stock solution was adjusted to pH 3 by the addition of concentrated HCl. An amount of 100 ml of the stock solution was placed in a beaker with a stirring bar. In separate experiments, the lanthanum oxycarbonates corresponding to Examples I, II and III of the present invention were added to the solution. The amount of lanthanum oxycarbonate or carbonate was such that the amount of La in suspension was 3 times the stoichiometric amount needed to react completely with the phosphate. Samples of the suspension were taken at time intervals through a filter that separated all solids from the liquid. The liquid samples were analyzed for phosphorous.

[0041] Two further experiments were run in the same conditions as those given in the previous paragraph, except that commercial lanthanum carbonate tetra hydrate $\text{La}_2(\text{CO}_3)_3 \cdot 4\text{H}_2\text{O}$ in one case, commercial lanthanum carbonate monohydrate $\text{La}_2(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$ in the other case, were added to the solution.

[0042] Curves showing the amount of phosphorous removed from the solution as a function of time with the different lanthanum compounds are given in FIG. 10. The figure shows that the rate of removal of phosphate with the different oxycarbonates of this invention is faster than the rate of removal obtained for commercial lithium carbonate tetra hydrate or monohydrate.

[0043] The particles of lanthanum oxycarbonate made according to the process of the present invention, particularly those made following the methods corresponding to Example II and Example III have the following common properties:

- They have low solubility in water.

- Their hollow shape gives them a high surface area, providing a fast reaction rate, while the particles themselves are aggregates large enough to be collected on ordinary water filters.
- They have faster phosphate binding kinetics than commercial grade lanthanum carbonates, as shown in FIG. 10.

5 **[0044]** Because of these characteristics, the products of the present invention have the potential to be used to remove phosphates from swimming pools and other water systems more efficiently than existing compositions and methods. Particularly, the products of the present invention have the potential of faster removal
10 of phosphates without forming small, unfilterable precipitate and without leaving unreacted La salts in solution, and to be used directly in the filtration system of a swimming pool. The oxycarbonate compounds are safe and do not need flocculants or ordinary chemicals. No pool downtime is needed to use them.

15 **[0045]** While the invention has been described in conjunction with specific embodiments, it is to be understood that many alternatives, modifications, and variations will be apparent to those skilled in the art in light of the foregoing description. Accordingly, this invention is intended to embrace all such alternatives, modifications, and variations that fall within the spirit and scope of the appended claims.